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Solution Stoichiometry Name Chem Worksheet 15-6. © John Erickson, 2005 WS15-6SolutionStoich. USEFUL EQUATIONS. molarity = . L solution mol solute. 1 L = 1000 mL. The molarity of a solution is a ratio of the moles of solute per liters of solution. The units for molarity are written as mol/L or M. This measurement is used to perform stoichiometric calculations.

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first . 1] How many grams of calcium phosphate can be produced from the reaction of 2.50 L of 0.250 M Calcium chloride with an excess of phosphoric acid?

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Answer Chem 15 2 Worksheets - Learny Kids

As we learned previously, double replacement reactions involve the reaction between ionic compounds in solution and, in the course of the reaction, the ions in the two reacting compounds are "switched" (they replace each other). Because these reactions occur in aqueous solution, we can use the concept of molarity to directly calculate the number of moles of reactants or products that will ...

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