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A. Buffer

solution B.

Neutral solution

C. Saturated

solution D.

Supersaturated

solution. 10.

Protein is a

polymer formed

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from the linkage
of A. Amino acid
molecules. B.
Fatty acid
molecules. C.
Glucose units.
D.

Monosaccharides.
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. SECTION A. 1.
Identify the
solid remaining
when each of the

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following is
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A solution of
0.20 mole of
NaBr and 0.20
mole of MgBr₂
in 2.0 dm³ of
water is to be
analysed. ...

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2 (ESSAY)

SECTION A. 1.

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Identify the solid remaining when each of the following is heated. ...

State whether R would be expected to form acidic or basic oxide. (b) (i) State two assumptions of the ...

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An aqueous solution that contains 1 mol (342 g) of sucrose in enough water to give a final volume of 1.00 L has a sucrose concentration of 1.00 mol/L or

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1.00 M. In
chemical
notation, square
brackets around
the name or
formula of the
solute represent
the molar
concentration of
a solute.
Therefore,
[sucrose] =
1.00M.

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A buffer solution (more precisely, pH buffer or hydrogen ion buffer) is an aqueous solution consisting of a mixture of a

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weak acid and its conjugate base, or vice versa. Its pH changes very little when a small amount of strong acid or base is added to it. Buffer solutions are used as a means of keeping pH at a nearly

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constant value
in a wide
variety of
chemical
applications.

Buffer solution
- Wikipedia

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best way to use
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formulas? Ans-
The best way to
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your textbook,
read the theory
given in the
textbook make
the notes of the
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definitions and
reactions given
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to calculate the
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